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Aromaticity Introduced by Antiferromagnetic Ligand Mediated Metal-Metal Interactions. Insights from the Induced Magnetic Response in $[\text{Cu}_6(\text{dmPz})_6(\text{OH})_6]$

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$[\text{trans-Cu}(\mu\text{-OH})(\mu\text{-dmpz})]_6$ (**1**), exhibits six Cu(II) centers effectively coupled through an ligand mediated mechanism leading to a diamagnetic ground state for a wide temperature interval. Here we investigate further magneto-structural correlations standing on the possible free electron precession along such a copper-based ring-like nanocoil mediated by the bridging ligands. We find that in **1**, the mediated antiferromagnetic coupling leads to the characteristic that reminisce the aromatic ring behavior through evaluation of both induced currents and shielding cones from relativistic density functional theory level. According to our calculations of gauge including magnetically induced current densities and the induced magnetic field, a sizable ring current strength susceptibility is obtained for the cyclic Cu-N-N-Cu and Cu-O-Cu pathways allowing a magnetic exchange between the copper centers. Our study suggests that $[\text{Cu}_6(\text{dmPz})_6(\text{OH})_6]$ consisting of an aromatic ring structure displays aromaticity and superexchange along the Cu-O-Cu and Cu-N-N-Cu backbones, which accounts for the 80 % and 20% of the overall ring current strength susceptibility, respectively. This unravels the presence of particular aromatic rings characteristics in coordination compounds without a direct metal-metal bond, where several formally paramagnetic centers are antiferromagnetically-coupled through supporting ligands. We envisage that our findings can be extended to other examples depicting ligand-mediated interaction between metal centers.

Introduction

Polynuclear metal coordination compounds represent an interesting class of molecules in the fields of inorganic chemistry and material science due to their intriguing structures, high connectivity and large stability.^{1–6} Such systems exhibit tailorable magnetic and catalytic properties, which are relevant for a wide range of technological applications.^{6–13} Since the early days of molecular magnetism, small complexes involving copper(II) ions have been the topic of intense research efforts by both synthetic and theoretical

chemists, focusing into the antiferromagnetic exchange coupling between the paramagnetic d^9 -Cu(II) ions.^{14–20} These joint efforts have led to novel classes of related molecules that function as valuable model systems to test new theories for understanding the role of the ligands connecting the metal centers.^{21–25}

The coordination chemistry of Cu(II) is a very active research field where the studies comprise synthesis of new polynuclear complexes by taking advantage of the structural versatility provided by N-donor ligands. The use of five-membered heterocyclic rings has shown to be effective in the synthesis of discrete metal-organic assemblies with variable nuclearities and sizes.^{26–32} Metal-organic structures involving six, eight, and larger number of copper atoms have been characterized, given by $[\text{trans-Cu}(\mu\text{-OH})(\mu\text{-dmpz})]_n$ ($n=6, 8$)^{33–35} and $[\text{cis-Cu}(\mu\text{-OH})(\mu\text{-pz})]_n$ ($n=6, 8, 9, 10$, etc.),³⁶ (pz= pyrazole; dmpz= 3,5-dimethylpyrazole). Moreover, complexes with high nuclearities have been also observed, leading to a unique and novel class of anion-sequestering agents (*nanojars*).^{37–39}

The $[\text{trans-Cu}(\mu\text{-OH})(\mu\text{-dmpz})]_6$ complex (**1**) involves six formally paramagnetic Cu(II) centers in a cyclic structure. However, a strong antiferromagnetic ground spin state ($S=0$) has been experimentally characterized via temperature

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dependent magnetic susceptibility measurements,³⁴ which is well separated from other spin states, ensuring the diamagnetic state.³⁴ This is also found for a related fluoromethyl derivative by Mohamed and coworkers.⁴⁰

Owing to the long Cu-Cu distances in **1** (3.181 Å), no direct orbital overlap within Cu₆ backbone is expected. Thus, the strong antiferromagnetic coupling observed between the paramagnetic centers is mediated by the ligands (superexchange interaction)^{22,24,25} being favored by the degree of covalency,⁴¹ which ensures the characterized overall singlet spin ground state. Hence, [*trans*-Cu(μ-OH)(μ-dmpz)]₆ complex appears as a suitable system for exploring the consequences of an effective ligand mediated antiferromagnetic coupling of cyclic metal-ligand structures.

The aim of this work is to investigate whether the ligand assisted Cu-Cu interactions in **1** can lead to an antiferromagnetic network favoring a free electron precession in such circular nanocoil when exposing the molecule to an external magnetic field. Since magnetically induced ring currents are typical for aromatic molecular rings,^{42,43} an eventual ring current relates the antiferromagnetic coupling and aromaticity, which reveals a novel property for cyclic molecular systems with ligand-mediated metal-metal interactions between paramagnetic centers. Thus, extending knowledge in magneto-structural relationships in the magnetochemistry field. The investigated structure exhibits two bonding backbones involving the hydroxyl- and pyrazole ligands, respectively, which are known to mediate the antiferromagnetic coupling between the Cu atoms and eventually also sustain magnetically induced ring currents.

Herein, we elucidate the magnetic behavior of the diamagnetic **1** in the presence of an external magnetic field. Two complementary approaches have been employed namely, calculations of the current density using the gauge including magnetically induced current (GIMIC) method^{43–45} and calculations of the magnetic response function in discrete points around the entire molecule. The current density calculations provide quantitative values for the current strength passing chosen planes in the molecule and it can also be used for visualizing the magnetically induced current density, which in turn give rise to the induced magnetic field that is obtained using the second approach.

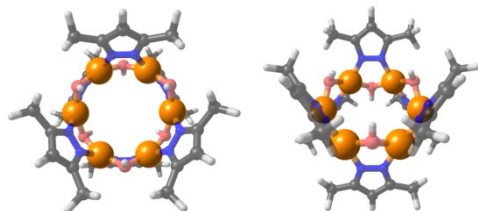


Figure 1. Two views of the optimized structure for [Cu₆(dmPz)₆(OH)₆]

Computational Details

Relativistic density functional theory (DFT) calculations⁴⁶ were carried out using the ADF code⁴⁷, incorporating scalar

relativistic effects via the one and two-component ZORA Hamiltonian⁴⁸. We employed all electron triple- ζ Slater basis sets augmented with two polarization functions (STO-TZ2P) for valence electrons, and the non-local Becke-Perdew (BP86) functional within the generalized gradient approximation (GGA).^{49–51} Geometry optimizations were performed without any symmetry constraints, via the analytical energy gradient method implemented by Versluis and Ziegler⁵². In order to consider long-range interactions, Grimme's dispersion correction was added for both geometry optimizations and energy decomposition analysis.⁵³

The nuclear magnetic shielding constants and nucleus-independent shielding tensors were calculated with the NMR module of ADF at the DFT BP86 level using the STO-TZ2P basis set employing gauge-including atomic orbitals (GIAO).^{54–57}

The magnetically induced current density susceptibilities were calculated using the gauge including magnetically induced current (GIMIC) method.^{43–45} In the GIMIC calculations, the input information comprises basis set data, the unperturbed and first order magnetically perturbed atomic orbital density matrices obtained from the Kohn-Sham and NMR calculations, respectively. The density matrices were obtained at the BP86/def2-TZVP level in combination with the resolution of the identity (RI) or density fitting approximation using Turbomole version 7.0.^{58–62}

A faster basis set convergence is achieved by using perturbation dependent basis sets for the calculation of the magnetically induced current density. Thus, current densities calculated with triple- ζ basis sets are close to the basis-set limit. Magnetically induced current density studies including an integration analysis of the current flow have proven to be a very reliable means for aromaticity assignments according to the magnetic criterion,^{43,45} in particular when investigating complex multiring molecules.^{45,63}

Aromaticity has been suggested to be multidimensional i.e., various means to measure it provides different degree of aromaticity.^{42,64–67} However, in more recent studies the multidimensionality of the aromaticity concept has been questioned,^{68,69} because for multiring systems aromaticity is a global property involving several rings along different pathways.⁴³ Many employed aromaticity criteria yield the degree of aromaticity of individual molecular rings,^{70–72} because they are not easily applicable to molecules consisting of many annelated rings.⁶³ Calculations of magnetically induced current strengths give a more accurate picture of the electron delocalization, since they show how and to what extent electrons move around the molecular rings when the molecule is exposed to an external magnetic field,⁴⁴ which is a relevant characteristic of aromatic species.^{42,66,73–75}

Results and Discussion

The molecular structure of [*trans*-Cu(μ-OH)(μ-dmpz)]₆ is shown in Figure 1. It has *D*_{3d} symmetry with the dmpz' and OH' ligands distributed in a *trans* arrangement, involving six formally Cu(II) centers. The calculated structure agrees with the experimental one,³⁴ suggesting that the molecular properties are well

accounted at the employed present level of theory. The Cu-Cu distances are 3.190 Å as compared with the experimental distance of 3.181 Å obtained in X-ray measurements. The calculated Cu-N and Cu-O distances are 1.952 and 1.948 Å, and the corresponding experimental values are 1.952 and 1.934 Å.³⁴

A molar magnetic susceptibility ($\chi_m T$) of 2.25 emu mol⁻¹ K is expected for six magnetically isolated copper(II) centers^{34,40}. In contrast, the smaller $\chi_m T$ product value experimentally measured of 0.5 emu mol⁻¹ K at room temperature, and close to zero at temperatures below 50 K, indicates that there is a strong antiferromagnetic coupling between the Cu centers³⁴ ensuring a diamagnetic S=0 ground state. From the broken-symmetry (BS) procedure developed by Noodleman^{76–78} a BS solution of the singlet ground state is obtained where each Cu center carries a ± 0.39 spin density in an alternate fashion (supporting information), accounting for the singlet state. A comparison between both restricted and BS states related to the overall diamagnetic ground state reveal that the latter is slightly favoured by 4.7 kcal/mol, suggesting that such solutions are comparable in terms of stability.

In this sense, the magnetic exchange between the copper ions can be transmitted through the Cu-N-N-Cu and Cu-O-Cu moieties with similar or different contribution to the overall antiferromagnetic coupling. In order to unravel the preferred magnetic coupling pathway, we evaluate the interaction energy of each ligand in **1**. The analysis of the interaction between Cu and the pyrazol or the hydroxyl group leading to the formation of **1** were carried for both [dmpz]⁻[Cu₆(dmpz)₅(OH)₆]⁺ and [OH]⁻[Cu₆(dmpz)₆(OH)₅]⁺ fragments, owing to their favorable oxidation state as isolated species. The charge distribution analysis of **1** based on Hirshfeld charges exhibits that the dmpz⁻ and OH⁻ ligands in the restricted state have charges of -0.49 and -0.53 a.u., respectively, and of -0.51 and -0.55 a.u. in the broken-symmetry state, as compared with their formal charge of -1 as isolated species. Thus the charge distribution in both solutions for the singlet ground state is similar. Owing that the detailed analysis of the bonding interaction is not available for broken-symmetry states, we based the following findings in the restricted state unless stated otherwise.

The obtained interaction energies (ΔE_{int}) for the bridging ligands are -163.0 and -169.5 kcal mol⁻¹, respectively indicating equal energy contributions from the Cu-OH-Cu and Cu-N-N-Cu moieties to the formation of the Cu₆ ring. A large metal-ligand orbital overlap connecting the six copper ions along the backbone is obtained, showing its important role for the stability of **1** and for the effective ligand-mediated spin coupling mechanism.⁴¹

To explore the nature of this interaction, we performed the energy decomposition analysis (EDA) within the Morokuma–Ziegler scheme^{79–81}. This allows describing ΔE_{int} in terms of different chemically meaningful contributing terms (see Table 1):

$$\Delta E_{\text{int}} = \Delta E_{\text{Pauli}} + \Delta E_{\text{orb}} + \Delta E_{\text{elstat}} + \Delta E_{\text{disp}}$$

The stabilizing ΔE_{elstat} term refers to the electrostatic character of the interaction, which is obtained by considering each defined fragment (e.g., A and B) in its unperturbed (frozen)

electron density as isolated species ($\Psi_A \Psi_B$). The repulsive ΔE_{Pauli} quantity accounts for the four-electron two-orbital interactions between occupied orbitals, which is calculated from the energy change due to the antisymmetrization and renormalization of the overlapping fragment densities ($\Psi_0 = N\hat{A}\{\Psi_A \Psi_B\}$). The stabilizing ΔE_{orb} term is obtained when the densities of the constituent fragments relax into the final molecular orbitals (Ψ_{AB}) accounting for the covalent character of the interaction. In addition, the pairwise dispersion correction⁵³ (DFT-D3) allows us to evaluate the dispersion interaction (ΔE_{disp}) related to London forces. To overcome basis set superposition errors (BSSE), the counterpoise correction was employed.⁸²

The ratio between the stabilizing quantities reported in Table 1, namely, ΔE_{orb} , ΔE_{elstat} and ΔE_{disp} , yields the overall electrostatic or covalent character of the metal-ligand interaction. For both ligands, the stabilizing nature of ΔE_{int} is mainly dominated by the ΔE_{elstat} term (~69%), which is the main contribution to the electrostatic character of the chemical bonding when forming **1** with formally Cu(II) centers and anionic ligands. ΔE_{elstat} are -284.3 and -278.5 kcal mol⁻¹, for the dmpz⁻ and OH⁻ ligands, suggesting a slightly more favorable electrostatic interaction in the former. The ΔE_{orb} terms, which are -115.7 and -125.8 kcal mol⁻¹, respectively, indicate that the charge-transfer interaction is also contributing significantly (~30%) to the overall electrostatic interaction of the Cu(II)-ligand bonding. The bond formation in the cyclic structure results from the overlap between Cu 3d orbitals and appropriate ligand orbitals, which in turn leads to the effective ligand-mediated spin coupling mechanism of the observed singlet spin ground state.³⁴

The different bonding contributions to the ΔE_{orb} term can be further described through the Energy Decomposition Analysis with Natural Orbitals for Chemical Valence^{83–85} (EDA-NOCV).

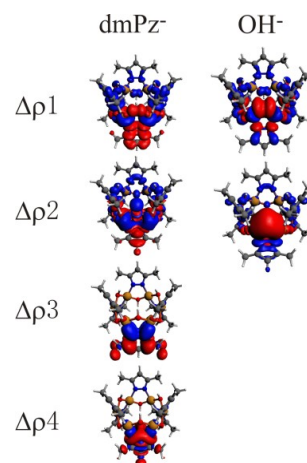


Figure 2. Deformation densities from the NOCV-EDA analysis. See text and Table 1.

When analyzing the major contributions to the orbital interaction term, as provided by the deformation densities of the NOCV analysis in Figure 2 and Table 1, it turns out that the bond formation along the structural backbones of **1** is of σ -character. The bonding interaction is accounted for by two main contributions to the deformation density, namely $\Delta\rho_1$ and

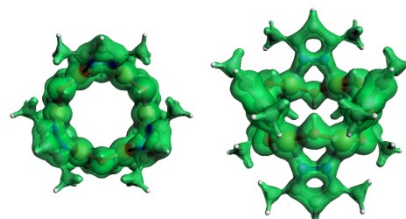
$\Delta\rho_2$, depicting the $\text{Cu} \leftarrow \text{dmpz}^-$ and $\text{Cu} \leftarrow \text{OH}^-$ charge transfers. The two major orbital interactions are sizable and contribute approximately 71% and 88% to the ΔE_{orb} term, compromising a charge estimation (Δq_n) of 0.80 and 0.90 a.u., respectively. The π -donation in dmpz^- covering the $\text{Cu} \rightarrow \text{dmpz}^-$ back bonding is given by $\Delta\rho_3$ and $\Delta\rho_4$, which contribute roughly 12% to the density difference (with a $\Delta q_3 = 0.28$ and $\Delta q_4 = 0.18$). The σ -contributions to the deformation densities ($\Delta\rho_1$ and $\Delta\rho_2$) obtained in both fragmentation schemes suggest that the bonding is delocalized over the structure involving both ligands.

Table 1. Energy decomposition analysis for ligand interaction, and the corresponding energy and charge transfer from each relevant density deformation channels from NOCV-EDA (ΔE_n^{orb} and Δq_n). The values for ΔE are given in $\text{kcal}\cdot\text{mol}^{-1}$ and Δq in a.u..

	dmpz ⁻		OH ⁻	
ΔE_{Pauli}	245.6		238.5	
ΔE_{Orb}	-115.7 (28.3%)		-125.8 (30.8%)	
ΔE_{Elstat}	-284.3 (69.6%)		-278.5 (68.3%)	
ΔE_{Disp}	-8.6 (2.1%)		-3.7 (0.9%)	
ΔE_{Int}	-163.0		-169.5	
	ΔE_n^{orb}	Δq_n	ΔE_n^{orb}	Δq_n
$\Delta\rho_1$	-45.9 (39.7%)	0.80	-70.0 (55.7%)	0.91
$\Delta\rho_2$	-36.7 (31.7%)	0.69	-41.1 (32.7%)	0.57
$\Delta\rho_3$	-6.1 (5.3%)	0.28		
$\Delta\rho_4$	-6.0 (5.2%)	0.18		

A comparison of the calculated values for ΔE_{orb} , shows that similar values are obtained for OH^- and dmpz^- , where OH^- most likely forms a more favorable overlap than dmpz^- . Both ligands offer a favorable overlap that connects the copper ions and thereby ensuring the characterized spin ground state, with a slightly larger contribution from OH^- . Thus, we obtain three different delocalized bonding situations supporting the Cu_6 ligand-mediated antiferromagnetic coupling in the overall diamagnetic structure of **1**. i) The magnetic coupling follows the $(\text{Cu-N-N})_6$ backbone involving all the copper ions and the six dmpz^- ligands in the *trans* arrangement. ii) Analogously, the $(\text{Cu-O})_6$ bonds of the OH^- groups couple the spins of the Cu atoms, or, iii) a combination of a coupling via both backbones.

Figure 3. Two views of the electron density involving the contribution from Cu 3d-orbitals (isosurface value = 0.03 a.u.).



See text.

To evaluate the magnetic coupling pathway between the six Cu atoms, we calculate the electron density for all molecular orbitals with significant contributions from the 3d orbitals of the Cu atoms. Contributions from 4s- and 4p-Cu orbitals were not observed. In this sense, we can determine the magnetic exchange pathway due to the metal-ligand overlap that leads to the singlet electronic state observed experimentally (*vide infra*). The electron density in Figure 3 shows that the 3d orbitals of the Cu atoms are delocalized along both $(\text{Cu-N-N})_6$ and $(\text{Cu-O})_6$ backbone leading to an efficient antiferromagnetic coupling between the six copper centers. Thus, the diamagnetic ground state of $[\text{trans-Cu}(\mu\text{-OH})(\mu\text{-dmpz})]_6$ is caused by the interaction between the Cu atoms and both ligands, which contribute in the same way to the Cu-Cu interaction.

The aromatic character of organic molecules has been widely interpreted in terms of the ring current contribution to magnetic shielding constants.^{70,75,87–89} Application of an external magnetic field (\mathbf{B}^{ext}) perpendicularly to a molecular ring leads to an induced field (\mathbf{B}^{ind}) opposed to the external one, resulting in a shielding response at the center of the ring. In contrast, for antiaromatic molecules, such induced magnetic field enhances \mathbf{B}^{ext} leading to deshielding inside the ring. Calculating \mathbf{B}^{ind} or actually the shielding tensor (σ) function⁵⁷ in discrete points in space around the studied molecules provides an overall picture of the short-ranged and long-ranged anisotropy cones, which are useful for understanding the magnetic anisotropies introduced by functional groups or molecules.^{90–92}

The shielding function due to the magnetically induced current density was calculated in selected planes in the molecule. The graphical representation of the magnetic response (\mathbf{B}^{ind}) of the applied external magnetic field (\mathbf{B}^{ext}) was obtained as $\mathbf{B}_i^{\text{ind}} = -\sigma_{ij} \mathbf{B}_j^{\text{ext}}$ by calculating the nucleus independent shielding tensor (σ_{ij}). Magnetic response is often given in terms of the more familiar magnetic chemical shift tensor ($\delta_{ij} = -\sigma_{ij}$). In the visualization of the induced magnetic field, the direction of the external magnetic field is chosen to be oriented along one of the Cartesian directions (*i*) showing the induced magnetic field along in other Cartesian directions (*j*). Analyses of given components of the shielding tensor provide an overall picture of the shielding and deshielding regions in the molecule.

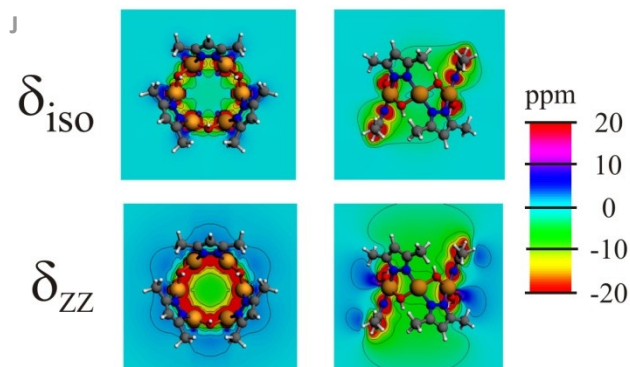


Figure 4. Shielding response under an applied field for **1**. Isotropic and zz-components are given.

The isotropic response given by δ_{iso} (Figure 4), accounts for the experimental molecular tumbling of molecules in solution and gas-phase NMR experiments. For the restricted case (see above) the calculated shielding function exhibits a slight overlap of the induced shielding cones between the five-membered aromatic rings (dmpz) towards the center of the structure. The isotropic contribution to the magnetic chemical shift in the center of the δ_{iso} Cu_6 ring, which is also called $\text{NICS}(0)$,^{71,72,88,89} is -2.20 ppm. $\text{NICS}(0)$ is the average of the response at the origin of **1** for the three different orientations of the external magnetic field. The shielding function shown in Figure 5 reach its maximum of -4.05 ppm at about 2.6 Å due to the shielding contributions originating from the pyrazole rings. For the broken-symmetry state, similar values are found denoting that a related response is obtained when the external field is applied with $\text{NICS}(0) = -2.15$ ppm and -4.09 ppm 2.6 Å at from the center.

The δ_{zz} component of the magnetic chemical shift function (also called NICS_{zz})^{77,78} provides information about the magnetic response in the z direction when the external magnetic field is applied in the same direction. Thus, assuming a simple ring-current picture, the δ_{zz} component is related to ring-current strength perpendicularly to the external magnetic field.^{70,94–96} The NICS_{zz} function reveals an interesting long-ranged shielding region at the center of **1**. The $\text{NICS}_{\text{zz}}(0)$ value at the center of the Cu_6 ring is $\delta_{\text{zz}} = -7.11$ ppm. Along the z-axis, the value of δ_{zz} reaches its maximum at the center of the molecule and decreases to -5.00 ppm at about 4.5 Å, and -3.0 ppm at 7.5 Å, illustrating the long range behavior of the shielding cone. A complementary deshielding region can also be observed. Similarly, for the broken-symmetry state the $\text{NICS}_{\text{zz}}(0)$ amounts to -7.34 ppm, with a shielding function profile strongly related to the restricted solution.

Thus, **1** in both restricted and broken-symmetry solutions shows the typical characteristics for aromatic rings in metal-ligand complexes. The obtained magnetic behavior is due to the metal-ligand orbital overlap of the zig-zag Cu-N-N-Cu and Cu-O-Cu backbones that connect the copper ions. The magnetic response suggests that the $[\text{trans-Cu}_6(\text{dmpz})_6(\text{OH})_6]$ complex is an aromatic molecule where the superexchange interaction of the molecular rings gives rise to its aromatic character.

Moreover, the theoretical evaluation of the $[\text{cis-Cu}_6(\text{Pz})_6(\text{OH})_6]$ counterpart, involving unsubstituted pyrazole rings, also

suggest a singlet spin ground state mediated by the ligands. Similarly to $[\text{trans-Cu}_6(\text{dmpz})_6(\text{OH})_6]$, a shielding response at the center of the structure ($\text{NICS}_{\text{zz}}(0) = -3.37$) is obtained when a magnetic field is perpendicularly applied in relation to the Cu_6 plane (supporting information). This support that the ligand-mediated Cu-Cu interactions induces an aromatic behavior to the overall structure.

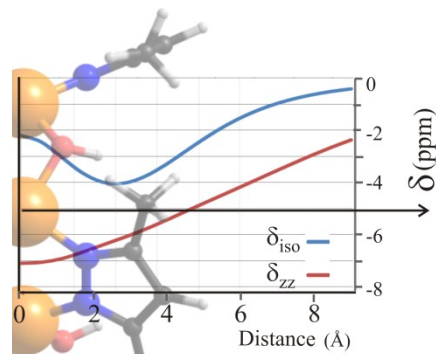


Figure 5. Shielding response profile along the z-axis. Isotropic and zz-components are given.

In order to evaluate an eventual electron precession along the ligand-mediated metal-metal interaction, which is typical for aromatic molecules,^{74,75} we calculated the current density induced by an external magnetic field (Figure 6) for the restricted case. Owing to the ring shape of **1**, the external field was oriented perpendicularly to the plane defined by the Cu_6 ring. Calculations of the current density using the GIMIC method combined with a numerical integration analysis of the ring-current strengths show that **1** sustains indeed a net diatropic ring current whose current-strength susceptibility is 9.9 nA/T. For comparison, the integrated ring-current strength susceptibility for benzene calculated at the same level of theory is 11.7 nA/T.⁸⁶ Thus, the molecular ring of **1** involving the six Cu atoms can be considered aromatic according to the obtained diatropic value for the magnetically induced ring current strength.

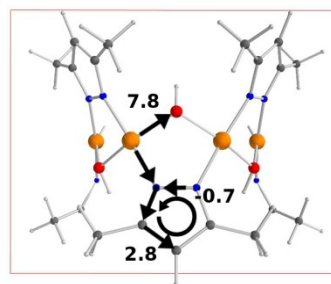


Figure 6. Calculated current pathways of Cu_6 . The numerical values are the integrated current strengths in nA/T calculated along the respective current pathway. Black arrows indicate the direction of the current flow assuming that diatropic currents circle counterclockwise. Copper atoms are displayed in orange, oxygen in red and nitrogen in blue, carbon atoms in black, and hydrogen atoms in light grey.

The induced current route along the $(\text{Cu-N-N})_6$ and $(\text{Cu-O})_6$ backbones (Figure 6), is reminiscent of the magnetic exchange

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pathways obtained from the analysis of the molecular orbitals with significant contributions from $3d$ orbitals of the Cu atoms. Thus, the magnetic exchange pathway and the electron current pathway are concordant. Bifurcation of the current was assessed by integrating the domains attributed to different pathways. The domains were chosen by visual inspection of the modulus of the ring current, as depicted in Figure 7. By comparing the strengths of the net ring current along the two backbones, one sees that the current splits at the Cu centers with 7.8 nA/T taking the Cu-O-Cu pathway and 2.1 nA/T is passing via the Cu-N-N-Cu route. The five-membered dmpz rings sustain a small local diamagnetic ring current of 0.7 nA/T, manifesting as an integrated net current strength of 2.8 nA/T for the dmpz⁻ ring. There is no through-space current between the Cu atoms, confirming the indirect Cu-Cu orbital interaction. Hence, the (Cu-O)₆ backbone sustaining 79% of the current strength is the preferred route for the magnetically induced current density accounting mainly for its aromatic character. About one fifth of the current strength passes along the (Cu-N-N)₆ backbone.

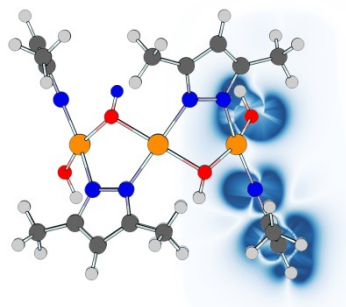


Figure 7. Modulus of the current density on a plane bisecting two copper atoms. Modulus of the current density was visualized with matplotlib using a cutoff value of 0.01 nA/T, and the plot was overlayed on isometrically visualized Cu₆ using GIMP.

Conclusions

The six paramagnetic Cu(II) centers forming a ring moiety with their ligands in $[trans-Cu(\mu-OH)(\mu-dmpz)]_6$, are antiferromagnetically coupled resulting in a diamagnetic ground state, as observed from experimental magnetic susceptibility measurements. The coupling is mediated by the bridging dimethylpyrazolate and hydroxyl ligands, which contribute to the magnetic exchange pathway through an effective metal-ligand overlap. The bonding between Cu₆ and the ligands is delocalized along the structure, which allows an electron precession when applying a magnetic field perpendicularly to the Cu₆ containing ring. Our study shows that a sizable ring-current strength susceptibility is obtained for the cyclic Cu-ligand structure indicating that the ring-shaped structure forms an aromatic ring. The Cu-O-Cu backbone is the preferred ring current pathway of the two possible routes accounting for the 80% of the overall induced current strength. $[Cu_6(dmpz)_6(OH)_6]$ forms an aromatic ring structure sustaining magnetically induced current flows along

both the Cu-N-N-Cu and the Cu-O-Cu backbones. Thus, the ligand-mediated spin coupling mechanism of cyclic metal-organic structures with magnetic exchange pathways such as in $[Cu_6(dmpz)_6(OH)_6]$ can be related to the aromatic character of its molecular rings. We envisage that the mechanism can be extended to comprise also other cyclic diamagnetic structures with individual paramagnetic centers that are magnetically coupled through organic ligands.

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